

Download File PDF Answers For Polyatomic Ions Packet

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*This is like a distributive in math. $2(x+3) \rightarrow 2x+6$
Everything inside of the parenthesis is multiplied by two.*

4. What does the subscript "2" inside the parentheses of the chemical formula tell you about the compound?
The subscript "2" outside of $\text{H}_2(\text{NO}_3)_2$ the parentheses indicate that there are two of the polyatomic ion NO_3^- (NO_3^-).

20. How many atoms of each element are in one formula unit of ammonium phosphate, $(\text{NH}_4)_3\text{PO}_4$?

nitrogen	hydrogen	phosphorus	oxygen
3	12	1	4

21. A student writes the chemical formula for the ionic compound calcium hydroxide as CaOH .

a. Write the chemical formula for each ion in the compound.
Calcium: Ca^{+2} Hydroxide: OH^{-1}

b. Why is the student's chemical formula for the compound calcium hydroxide wrong?
The subscript of "2" after hydrogen (H) indicates that hydrogens are not being balanced and just one ion is present. It should be placed outside of the parenthesis and a subscript of "2" should be placed outside of the parenthesis.

22. Many of the chemical formulas in Model 3 include parentheses. Which one of the following statements is the appropriate use of parentheses in binary ionic compounds? For the three rules, provide an example that does not apply in all cases, show at least one counter example from the chemical formulas in Model 3.

*Not true - $\text{Al}_2(\text{CO}_3)_3$ - Al has more than one atom.
Not true - KMnO_4 - Mn has no parentheses because there is only one.
Not true - NH_4^+ - has no parentheses because there is only one.*

23. Parentheses are used around any polyatomic anions used more than once in a formula unit.
Not true - $(\text{NH}_4)_3\text{PO}_4$ - ammonium is a polyatomic cation.

24. Parentheses are used around any polyatomic ions.

25. Write chemical formulas for the following binary ionic compounds.

a. Calcium sulfide	b. Copper(II) nitrate	c. Lithium phosphate
$\text{Ca}^{+2}\text{S}^{2-} \rightarrow \text{CaS}$	$\text{Cu}^{+2}\text{NO}_3^- \rightarrow \text{Cu}(\text{NO}_3)_2$	$\text{Li}^{+}\text{PO}_4^{3-} \rightarrow \text{Li}_3\text{PO}_4$
d. Potassium permanganate	e. Aluminum sulfide	f. Magnesium chloride
$\text{K}^{+}\text{MnO}_4^- \rightarrow \text{KMnO}_4$	$\text{Al}^{+3}\text{S}^{2-} \rightarrow \text{Al}_2(\text{S}_3)_3$	$\text{Mg}^{+2}\text{Cl}^- \rightarrow \text{MgCl}_2$

Polyatomic Ions 5 Key

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