

#Jenny



Finally I get this ebook, thanks for all these I can get now!

#Rio



Cool! I'am really happy

#Markus Jensen



I did not think that this would work, my best friend showed me this website, and it does! I get my most wanted eBook

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#Che Salsa



My friends are so mad that they do not know how I have all the high quality ebook which they do not!

#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

101 - Chemistry Name: _____

Balancing Redox Equations

Syllabus reference:

"Write ionic half-equations and use them to construct full ionic equations."

Aim of this worksheet:

To introduce you to the idea of balancing redox equations by adjusting coefficients and by adding OH⁻, H₂O, and H⁺ where necessary

Task: Balance the following redox reactions:

- $\text{SO}_2(\text{g}) + \text{HNO}_3(\text{aq}) \rightarrow \text{H}_2\text{SO}_4(\text{aq}) + \text{NO}(\text{g})$
- $\text{Al}(\text{s}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{Al}_2(\text{SO}_4)_3(\text{aq}) + \text{H}_2(\text{g})$
- $\text{Au}^{3+}(\text{aq}) + \text{I}^{-}(\text{aq}) \rightarrow \text{Au}(\text{s}) + \text{I}_2(\text{s})$
- $\text{S}^{2-}(\text{aq}) + \text{I}_2(\text{s}) \rightarrow \text{SO}_4^{2-}(\text{aq}) + \text{I}^{-}(\text{aq})$
- $\text{Br}_2(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightarrow \text{Br}^{-}(\text{aq}) + \text{BrO}_2^{-}(\text{aq})$
- $\text{H}_2\text{O}_2(\text{aq}) + \text{ClO}_2^{-}(\text{aq}) \rightarrow \text{O}_2(\text{g}) + \text{ClO}_2^{-}(\text{aq})$
- $\text{Mn}(\text{s}) + \text{HNO}_3(\text{aq}) \rightarrow \text{Mn}^{2+}(\text{aq}) + \text{NO}_2(\text{g})$
- $\text{I}_2(\text{s}) + \text{OCl}^{-}(\text{aq}) \rightarrow \text{IO}_3^{-}(\text{aq}) + \text{Cl}^{-}(\text{aq})$
- $\text{Cr}_2\text{O}_7^{2-}(\text{aq}) + \text{HNO}_2(\text{aq}) \rightarrow \text{Cr}^{3+}(\text{aq}) + \text{NO}_3^{-}(\text{aq})$
- $\text{CrO}_4^{2-}(\text{aq}) + \text{S}^{2-}(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s}) + \text{S}(\text{s})$

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