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Name _____ Class _____ Date _____

Assessment

Chapter 8 Practice Test

Chapter: Chemical Equations and Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

1. You mix solution A with solution B in a beaker. Which of the following observations does not help you prove that a chemical reaction has occurred?
 - a. The beaker becomes warm.
 - b. You see gas bubble out of solution.
 - c. Solution A, solution B, and the mixture are all yellow.
 - d. None of the above; all help prove that a chemical reaction has occurred.
2. In the reaction described by the word equation sodium + water \rightarrow sodium hydroxide + hydrogen gas
 - a. sodium hydroxide is a product.
 - b. sodium is a product.
 - c. water is a reactant.
 - d. Both (a) and (c)
3. Which of the following indicates that a chemical equation is balanced?
 - a. The numbers of atoms of each element are the same on both sides of the equation.
 - b. All of the coefficients are the same.
 - c. The numbers of molecules on each side of the equation are equal.
 - d. The sums of the coefficients on each side of the equation are equal.
4. From a complete and correctly written chemical equation, you can obtain the
 - a. chemical formulas of the reactants and products.
 - b. relative amounts of the reactants and products.
 - c. physical states of the reactants and products.
 - d. All of the above
5. To balance a chemical equation, you must
 - a. adjust the subscripts.
 - b. adjust the number of products.
 - c. adjust the number of reactants.
 - d. adjust the coefficients.

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