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so many fake sites. this is the first one which worked! Many thanks

Example 4 Solution: The Equilibrium Constant Expression with Partial Pressures

The equilibrium  $\text{NO}_2$  partial pressure is 1.26 atm and  $\text{N}_2\text{O}_4$  is 0.199 atm:

$$2 \text{NO}_2 \rightleftharpoons \text{N}_2\text{O}_4$$
$$K = \frac{P_{\text{N}_2\text{O}_4}}{P_{\text{NO}_2}^2} = \frac{0.199}{1.26^2} = 0.125$$

NOTE: These are not the same experimental conditions from the previous problem.

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