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Gpb Chemistry Answers

Atomic Structure Chart Worksheet Name Key
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1) Define the following terms
Atomic Number: number of protons
Mass Number: sum of protons and neutrons
Atomic Mass: average mass of an atom
Isotope: atoms of the same element with different numbers of neutrons
Isotope Name:
Nuclear Symbol:

2) Give the nuclear symbol for an atom with 9 protons and 10 neutrons: $^{19}_9\text{F}$
3) Give the nuclear symbol for an atom with atomic number 72 and 100 neutrons: $^{172}_{72}\text{Hf}$
4) Write the isotope name for an element with 20 protons and 21 neutrons: $^{41}_{20}\text{Ca}$

5) Complete this chart for the following NEUTRAL atoms using your periodic table:

Element	Atomic #	Protons	Electrons	Neutrons	Mass Number	Isotope Name	Nuclear Symbol
Si	14	14	14	15	29	Si-29	$^{29}_{14}\text{Si}$
B	5	5	5	5	10	B-10	$^{10}_5\text{B}$
F	9	9	9	10	19	F-19	$^{19}_9\text{F}$
U	92	92	92	143	235	U-235	$^{235}_{92}\text{U}$
Kr	36	36	36	48	84	Kr-84	$^{84}_{36}\text{Kr}$
U	92	92	92	146	238	U-238	$^{238}_{92}\text{U}$
Ag	47	47	47	71	118	Ag-118	$^{118}_{47}\text{Ag}$

6) Which two of the above would be an example of isotopes?
U-235 & U-238

7) Explain the difference between "atomic mass" and "mass number".
Mass number refers to a specific isotope of an element. Atomic mass is an average of all isotopes.

8) Complete this chart.

Particle	Charge	Mass in amu	Location in Atom	Discovered by
proton	+	1	nucleus	Rutherford
neutron	0	1	nucleus	Chadwick
electron	-	1/1836	outside nucleus	Thomson